

- Find the molar mass of the following compounds
 - $\text{Zn}(\text{CH}_3\text{COO})_2$
 - $\text{Fe}_2(\text{SO}_3)_3$
 - $\text{Mg}_3(\text{PO}_4)_2$
 - $\text{Na}_2\text{C}_2\text{O}_4$
 - $(\text{NH}_4)_3\text{AsO}_4$
- A chemical reaction calls for 0.250 moles of table salt (sodium chloride). How many grams of table salt are needed? (*Answer = 14.6 g*)
- A chemical reaction calls for 0.750 moles of potassium chloride. How many grams of potassium chloride are needed?
- How many moles of ammonium nitrate are in 29.63 grams? (*Answer = 0.3700 mol*)
- How many moles of lithium sulfate are in 35.45 grams?
- Suppose we had to convert 10.0 grams of Na_2SO_4 into moles. How many moles would there be? (*Answer = 0.0704 mol Na_2SO_4*)
- How many moles of iron (III) oxide are contained in 92.2 g? (*Answer = 0.577 mol*)
- Determine the number of moles of each of the following
 - 100 g of KMnO_4
 - 74 g of KCl
- Determine the number of grams of each of the following
 - 1.70 moles of KMnO_4
 - 0.25 moles of KCl