

**Chemistry 112****VSEPR Theory Questions**

Read through the slides on VSEPR Theory to answer the questions. In short, VSEPR Theory is the shapes molecules take on based on their lone pairs and bonding pairs.

1. What does VSEPR stand for?
2. How is it assumed that molecules achieve the most stable condition?
3. What are the 5 shapes that molecules can make according to VSEPR Theory?
4. Which groups can have central atoms?
5. Create the following chart in your book and, using the following elements, fill in the table.



Molecular Compound (molecular formula + structural diagram)	Shape formed by the molecule	How many shared/bonded pairs are there?	How many unshared/ lone pairs are there?	What is the Central Atom and what Group is it in?