

Science 10 – Chemistry Review for Test #2

1. What is the difference between:
 - a. Ionic, Polyatomic, and Molecular Compounds
 - b. Polyatomic Ion and Polyatomic Compound
 - c. Electron, Neutron and Proton
 - d. Skeleton equation and a word equation

2. Complete the following table below:

Element	Atomic Mass	Atomic Number	Protons	Neutrons	Electrons
		14			
	80				
hydrogen					
			7		
				24	
					28

3. Answer the following questions. If the symbols are given, find the name. If the name is given, find the chemical formula.

a. IONIC COMPOUND

- i. LiF
- ii. calcium nitride
- iii. nickel (III) phosphide
- iv. HgBr₂
- v. hydrogen chloride
- vi. tin (IV) oxide
- vii. sodium phosphide
- viii. Sc₂S₃
- ix. mercury (I) bromide
- x. barium nitride

b. POLYATOMIC COMPOUND

- i. BeSO_4
- ii. potassium chlorate
- iii. magnesium hydroxide
- iv. $\text{Ba}(\text{NO}_3)_2$
- v. calcium silicate
- vi. hydrogen nitrate
- vii. tin (II) sulfite
- viii. LiOH
- ix. $\text{Ni}_2(\text{SO}_3)_3$
- x. sodium oxalate

c. MOLECULAR COMPOUNDS

- i. C_3H_6
- ii. NO_2
- iii. tetranitrogen heptaoxide
- iv. dicarbon pentaoxide
- v. P_4Cl
- vi. octacarbon nonobromide
- vii. decanitrogen pentaiodide
- viii. dihydrogen monoxide
- ix. heptaselenium dichloride
- x. S_2Br_3

4. What is the Law of Conservation of mass?
5. Give an example of each of the following terms
 - a. Subscript
 - b. Coefficient
 - c. Word equation
 - d. Skeleton equation
6. There are several questions involving balancing chemical equations that I have given out before as handouts. Use those as practice. The answers have been posted for the first set of balancing equation problems.