

Solubilities of Ionic Compounds in Water		
Compounds	Solubility	Exceptions
Salts of Group 1A metals and ammonia	Soluble	Some lithium compounds
Ethanoates, nitrates, chlorates, and perchlorates	Soluble	Few exceptions
Sulfates	Soluble	Compounds of Pb, Ag, Hg, Ba, Sr, and Ca
Chlorides, bromides, and iodides	Soluble	Compounds of Ag and some compounds of Hg and Pb
Sulfides and hydroxides	Most are insoluble	Alkali metal sulfides and hydroxides are soluble. Compounds of Ba, Sr, and Ca are slightly soluble.
Carbonates, phosphates, and sulfites	Insoluble	Compounds of the alkali metals and of ammonium ions

Solubility Product Constants (K_{sp}) at 25°C					
Salt	K_{sp}	Salt	K_{sp}	Salt	K_{sp}
Halides		Sulfates		Hydroxides	
AgCl	1.8×10^{-10}	PbSO ₄	6.3×10^{-7}	Al(OH) ₃	3.0×10^{-34}
AgBr	5.0×10^{-13}	BaSO ₄	1.1×10^{-10}	Zn(OH) ₂	3.0×10^{-16}
AgI	8.3×10^{-17}	CaSO ₄	2.4×10^{-5}	Ca(OH) ₂	6.5×10^{-6}
PbCl ₂	1.7×10^{-5}	Sulfides		Mg(OH) ₂	7.1×10^{-12}
PbBr ₂	2.1×10^{-6}	NiS	4.0×10^{-20}	Fe(OH) ₂	7.9×10^{-16}
PbI ₂	7.9×10^{-9}	CuS	8.0×10^{-37}	Carbonates	
PbF ₂	3.6×10^{-8}	Ag ₂ S	8.0×10^{-51}	CaCO ₃	4.5×10^{-9}
CaF ₂	3.9×10^{-11}	ZnS	3.0×10^{-23}	SrCO ₃	9.3×10^{-10}
Chromates		FeS	8.0×10^{-19}	ZnCO ₃	1.0×10^{-10}
PbCrO ₄	1.8×10^{-14}	CdS	1.0×10^{-27}	Ag ₂ CO ₃	8.1×10^{-12}
Ag ₂ CrO ₄	1.2×10^{-12}	PbS	3.0×10^{-28}	BaCO ₃	5.0×10^{-9}